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Metallic Bonding ...

Section 7.3 – Bonding in Metals. The valence electrons of metal atoms can be modeled as a sea of electrons. Metallic bonds consist of the attraction of the free-floating valence electrons for the positively charged metal ions. Metals are good conductors and malleable because of their mobile electrons.

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Chapter 7 - Ionic and Metallic Bonding

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7.1 Ions - 7.1 Lesson Check - Page 199:
1 Answer To find the number of valence
electrons in an a representative element
simply look at the group number.

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In forming ionic bonds, atoms tend to attain the: A state of higher energy; The electron configuration of noble gas atoms; The electron configuration of halogen atoms; All of the above. A metallic bond is a bond between: Valence electrons and positively charged metal ions; The ions of two different metals; A metal and nonmetal; None of

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Chapter 7 "Ionic and Metallic Bonding ...
7.2 - Ionic Bonds and Ionic Compounds.

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Ionic Bonding. Anions and cations are held together by . opposite. charges (+ and -) Ionic compounds are called . salts. Formula unit - Simplest . ratio. of elements in an ionic compound.

Chapter 7 Ionic and Metallic Bonding

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Bonding, Review DRAFT. 9th - 12th grade. 68 times. Chemistry. 52% average accuracy. 3 years ago. kirch. 0. Save. Edit. ... Which element when combined with chlorine would most likely form an ionic compound? answer choices . Lithium. Carbon. Phosphorus. Bromine. Tags: Question 21 . SURVEY . 30 seconds . Q ...

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Lesson 7.1 Reading and Study Workbook

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Ionic and Metallic Bonding. BONDING AND INTERACTIONS. 7.1 Ions.

Essential Understanding Ions form when atoms gain or lose valence electrons, becoming electrically charged. Lesson Summary. Valence Electrons Valence

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electrons are the electrons in the outermost occupied energy level and are involved in ion formation. For a representative element, the group number equals the ...

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Noble gases are excluded from Table 6.2 (Electronegativity values) because they do not form many compounds, or in other words, they rarely react.

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ionic bond is the electrostatic attraction that binds oppositely charged ions together.

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 Present your graph to the group. Describe the general trend and any notable exceptions. How does

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this differ from the situation in ionic compounds? Explain the relationship between the properties of an element and the number of valence electrons that it contains. Write electron configurations to show the first two ionization processes for potassium. Answer and Explanation ...

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